If the reaction shown uses an excess of Br_2 and goes to completion, a total of how many bromine atoms will be in the final product?

A) 0

$$\begin{array}{c} \text{Br}_2 \\ \text{CH}_3 \end{array} \xrightarrow{\text{(excess)}} \\ \text{NaOH} \end{array} \qquad \begin{array}{c} \text{B) 1} \\ \text{C) 2} \\ \text{D) 3} \\ \text{E) 4} \end{array}$$