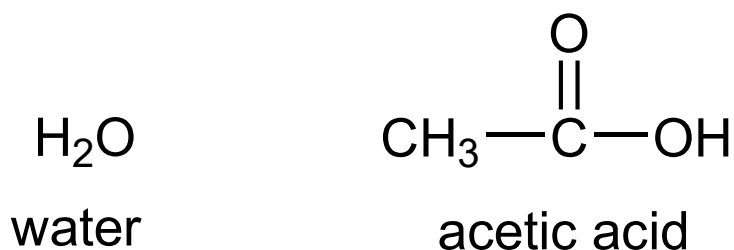


Which is the stronger acid? Explain briefly.



A) Because this is more stable: $\text{H}-\text{O}^{\ominus}$
water is the **stronger** acid.

B) Because this is more stable: $\text{H}-\text{O}-\text{H}$
water is the **weaker** acid.

C) Because this is more stable: $\text{CH}_3-\overset{\text{O}}{\parallel}{\text{C}}-\text{O}^{\ominus}$
acetic acid is the **stronger** acid.

D) Because this is more stable: $\text{CH}_3-\overset{\text{O}}{\parallel}{\text{C}}-\text{OH}$
acetic acid is the **weaker** acid.

E) It's impossible to predict acid strength without $\text{p}K_{\text{a}}$ data.