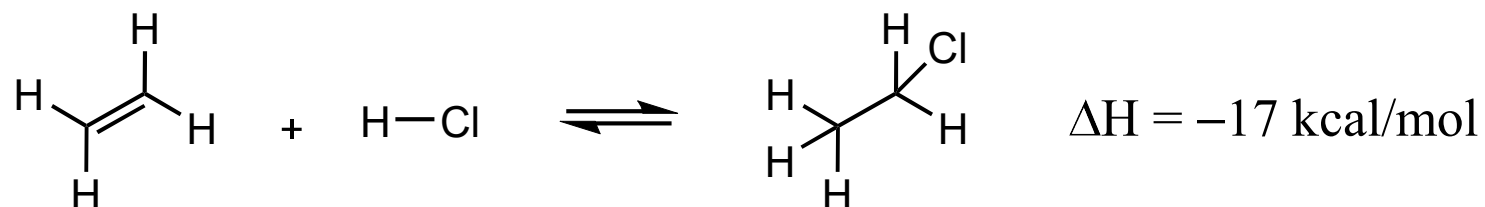


Do you expect the forward reaction to be spontaneous?
Explain briefly.



- A) The forward reaction will be spontaneous at all temperatures because it is exothermic and there is an increase in entropy.
- B) The forward reaction will be spontaneous at low temperatures because it is exothermic and there is an increase in entropy.
- C) The forward reaction will be spontaneous at high temperatures because it is exothermic and there is an increase in entropy.
- D) The forward reaction will be spontaneous at low temperatures because it is exothermic and there is a decrease in entropy.
- E) The forward reaction will be spontaneous at high temperatures because it is exothermic and there is a decrease in entropy.