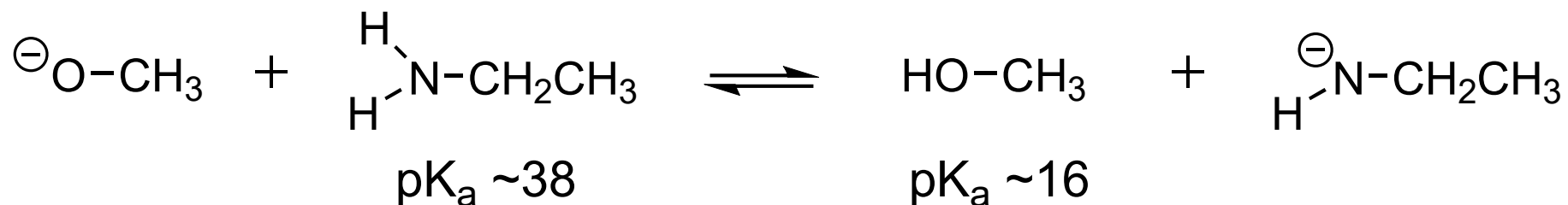


Use the given pK_a values to determine the direction of the equilibrium (forward or reverse favored?). Explain briefly.



- A) **Reverse** reaction is favored, because $\text{H}_2\text{NCH}_2\text{CH}_3$ is the stronger acid.
- B) **Reverse** reaction is favored, because $\text{H}_2\text{NCH}_2\text{CH}_3$ is the weaker acid.
- C) **Forward** reaction is favored, because $\text{H}_2\text{NCH}_2\text{CH}_3$ is the stronger acid.
- D) **Forward** reaction is favored, because $\text{H}_2\text{NCH}_2\text{CH}_3$ is the weaker acid.