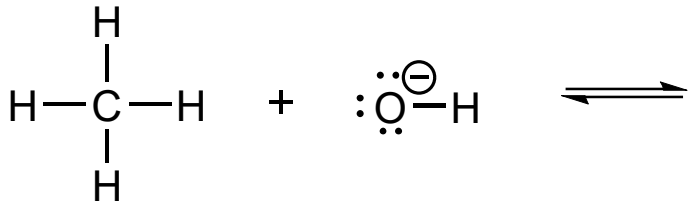
Is hydroxide a strong enough base to deprotonate methane (CH_4)? Explain.



A) Because hydroxide is more stable than $H_3C^{\ominus}$, hydroxide is a suitable base to deprotonate methane.

B) Because hydroxide is less stable than $H_3C^{\ominus}$, hydroxide is a suitable base to deprotonate methane.

C) Because hydroxide is more stable than $H_3C^{\ominus}$, hydroxide is NOT a suitable base to deprotonate methane.

D) Because hydroxide is less stable than $H_3C^{\ominus}$, hydroxide is NOT a suitable base to deprotonate methane.

E) It's impossible to predict the direction of the equilibrium without pK_a data.