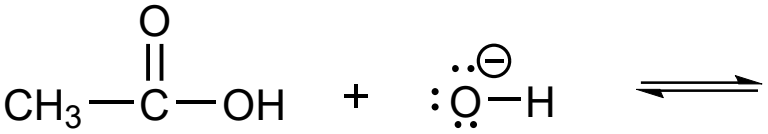
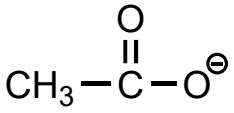
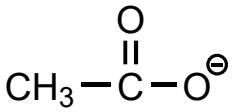
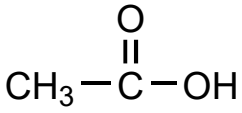
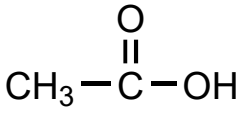
Is hydroxide a strong enough base to deprotonate acetic acid (CH_3CO_2H)? Explain.



A) Because hydroxide is less stable than this:  hydroxide is a suitable base to deprotonate acetic acid.

B) Because hydroxide is less stable than this:  hydroxide is NOT a suitable base to deprotonate acetic acid.

C) Because hydroxide is less stable than this:  hydroxide is a suitable base to deprotonate acetic acid.

D) Because hydroxide is less stable than this:  hydroxide is NOT a suitable base to deprotonate acetic acid.

E) It's impossible to predict the direction of the equilibrium without pK_a data.