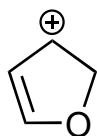
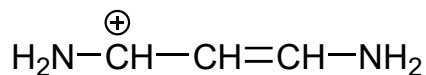
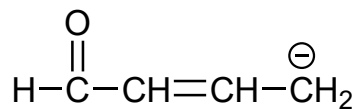
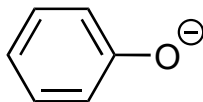
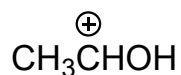
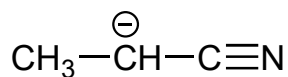
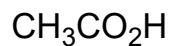


## Lewis Structures &amp; Resonance

Name: \_\_\_\_\_ Section: \_\_\_\_\_ (day/time)

**STOP! First work on textbook problems in Sections 2.10 and 2.11, including SkillBuilder 2.7!**For each of the following compounds/ions, **draw the significant resonance forms**. Be sure to draw complete Lewis structures (show all lone pairs), and use curved arrows to convert one drawing to the next.**Rank each structure** in terms of its contribution to the resonance hybrid, and briefly **explain** your rankings.**NOTE:** if a structure is charged then the goal is to find resonance structures that move the charge, since spreading a charge among multiple atoms ("delocalizing" the charge) will stabilize the species.*hint: first add in the missing H's*